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are related to  
hydrogen ion  
( $\text{H}^+$ )  
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experiment.  $\text{Ni}(\text{NO}_3)_2$   
+  $\text{H}_2\text{O}$ .  $\text{Ni}^{2+}$  reacts  
with concentrated  
ammonia s....  $\text{Ni}^{2+}$   
reacts with ethylene  
diamine  $\text{Ni}^{2+}$ ....

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 $\text{Cu(OH)}_2 + 2 \text{NH}_3 + 2$   
 $\text{NH}_4^+ \cdot [\text{Ni(H}_2\text{O)}_6]^{2+} \dots$   
 $\text{Ni}^{2+} + 2 \text{NO}_3^- \dots$   
hexaaquanickel (II)...  
green coloured....

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goal is to determine the molar concentration of two acid solutions by conducting titrations with a base of known concentration. You will be testing a strong acid, HCl, solution and a weak acid, HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>, solution.

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